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Impacts of Repeated Redox Cycling on Technetium Mobility in the Environment

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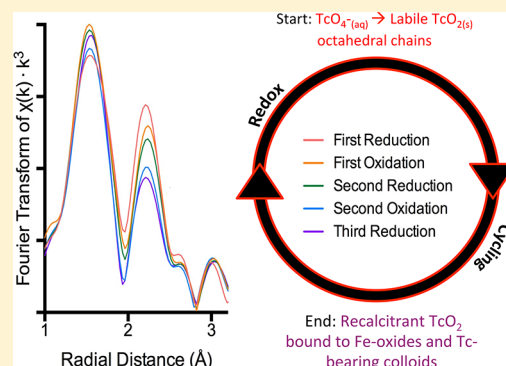
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S Supporting Information

ABSTRACT: Technetium is a problematic contaminant at nuclear sites and little is known about how repeated microbiologically mediated redox cycling impacts its fate in the environment. We explore this question in sediments representative of the Sellafield Ltd. site, UK, over multiple reduction and oxidation cycles spanning ~1.5 years. We found the amount of Tc remobilised from the sediment into solution significantly decreased after repeated redox cycles. X-ray Absorption Spectroscopy (XAS) confirmed that sediment bound Tc was present as hydrous TcO₂-like chains throughout experimentation and that Tc's increased resistance to remobilization (via reoxidation to soluble TcO₄[−]) resulted from both shortening of TcO₂ chains during redox cycling and association of Tc(IV) with Fe phases in the sediment. We also observed that Tc(IV) remaining in solution during bioreduction was likely associated with colloidal magnetite nanoparticles. These findings highlight crucial links between Tc and Fe biogeochemical cycles that have significant implications for Tc's long-term environmental mobility, especially under ephemeral redox conditions.



INTRODUCTION

The geological disposal of radioactive wastes and the safe management of radionuclide-contaminated land are global environmental priorities of intense public concern, which require comprehensive knowledge of the long-term behavior of radionuclides in the geosphere. Technetium-99 (⁹⁹Tc) is a β -emitting, high-yield fission product, which is long-lived (2.1×10^5 years) and environmentally persistent in the subsurface at nuclear mega-sites.^{1–5} Technetium is also present in radioactive wastes destined for geological disposal and has a high solubility, bioavailability, and bioaccumulation potential. It is thus a critical risk-driving radionuclide.^{4–8}

The solubility of Tc in the environment is largely governed by its oxidation state. Highly soluble pertechnetate (TcO₄[−]) dominates in oxic groundwater, sorbing weakly to geosolids (e.g.,^{9–11}). In contrast, poorly soluble Tc(IV) dominates under anoxic conditions, forming Tc-oxides, -sulfides, or sorbing strongly to mineral phases (e.g., Fe-oxides and sulfide minerals).^{11–13} In many environments, reductive scavenging of Tc(VII) to poorly soluble Tc(IV) coincides with the development of microbially mediated Fe(III)- and sulfate-reducing conditions. As a consequence, and in an effort to limit Tc migration in the environment, there has been widespread

interest in promoting Tc(VII) bioreduction in the shallow subsurface. For example, an electron donor such as acetate is added to groundwater to stimulate microbially mediated anaerobic processes.^{8,11,14–17} Electron donor release from radioactive waste packages in a geological disposal facility may also stimulate microbial reduction¹⁸ and this newly recognized “biobarrier” may serve to limit radionuclide migration through the deep subsurface.^{19,20}

To date, work in pure culture, sediment microcosms, flow-through columns, and in the field, has documented a range of Tc(VII) reduction mechanisms and Tc(IV) reduction products. Enzymatic Tc(VII) reduction can be facilitated by a range of microorganisms but is thought to be only modestly important in controlling Tc fate under typical environmental conditions.^{21,22} Instead, Tc(VII) reduction is dominated by reaction with byproducts of microbial metabolism, principally biogenic Fe(II) and sulfide.^{10,23–25} This includes abiotic reduction by Fe(II) in a range of Fe(II)-bearing minerals (e.g., refs

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17,20,26–29), as well as sulfide phases.^{12,16,30} Under Fe(III)-reducing conditions, and at Tc concentrations above the predicted solubility threshold of the TcO_2 species expected, Tc(IV) reaction products include TcO_2 octahedra/octahedral chains and/or discrete hydrous TcO_2 -phases that may sorb to Fe-(oxyhydr)oxides.^{13,20,26,27,31,32} Tc(IV) can also be incorporated into magnetite²⁹ and goethite,³³ and TcS_x -like phases have also been observed.^{12,16,34} Further, under these conditions, Tc sorbs strongly to sediments, presumably as Tc(IV), even below the predicted TcO_2 solubility threshold.^{11,35,36}

Despite the potential for removal of Tc from groundwater as poorly soluble Tc(IV), it is clear that with such a long half-life, understanding ^{99}Tc behavior over extended time scales is critically important. For example, reoxidation of Tc(IV) containing materials may result in Tc remobilization as soluble Tc(VII), leading to migration. Oxidizing conditions in the shallow subsurface can occur via groundwater perturbations or flooding, while over geological time scales, the introduction of oxic groundwater to a geological disposal facility is also of concern.^{37–40} To date, work exploring the stability of Tc(IV) reduction products over extended time scales has been relatively sparse, with most studies focusing on short duration (days to several weeks) reoxidation scenarios, typically involving exposure of reduced, Tc(IV)-bearing sediments to O_2 or NO_3^- rich groundwater (e.g.,^{26,32,37,39,41}). In this past work, significant (20–80%) Tc remobilization to solution has been reported, but in most cases some Tc remains as Tc(IV) within the reoxidized sediment.^{32,33,37,39,41,42} Binding of short-chain TcO_2 to Fe-(oxyhydr)oxides has been suggested as a mechanism that limits oxidative dissolution.³³ Similar studies with Tc-labeled synthetic Fe(II)-bearing mineral systems have also shown Tc(IV) retention when the minerals are oxidized.^{12,29,31,33} Here, incorporation of Tc(IV) into mineral lattice sites, physical shielding of Tc(IV) by Fe(III) oxidation rinds, and/or oxidative buffering by residual Fe(II) have all been suggested as mechanisms which could protect Tc(IV) from oxidative remobilization.^{26,29,31}

These studies, conducted over one redox cycle and for relatively short amounts of time (nominally days to several weeks), provide insights into Tc biogeochemical behavior in sediment and mineral systems highlighting that pseudo-irreversible binding of a fraction of Tc(IV) may be possible. However, ^{99}Tc has a long half-life and will likely undergo repeated redox cycling in the natural environment and the impact of such processes has not been investigated. Here, we investigate the hypothesis that multiple redox cycling events can impact on Tc solubility due to molecular scale changes in its speciation and fate. The objectives of this study were to define ^{99}Tc reactivity and fate across repeated redox cycles in nuclear legacy relevant sediments and to define the molecular scale speciation of ^{99}Tc during redox cycling to link these to macro scale remobilization behavior.

Our experiments used acetate as an electron donor to promote microbial reduction pathways in a model sediment system, representative of the Sellafield Ltd. site subsurface that was amended with $\text{Tc(VII)}_{(\text{aq})}$. This led to reduction of Tc(VII) and creation of $\text{Tc(IV)}_{(\text{s})}$ -phases. The reduced sediment was then oxidized (via addition of O_2) and reduced (via acetate addition) several times. Using sequential extractions, XAS, and TEM, our results detail an increased recalcitrance of sediment-bound Tc(IV) to oxidative remobilization across multiple redox cycles. This suggests a pathway to enhanced Tc(IV) retention in sediment systems; a positive

prospect for both contaminated land and radioactive waste geological disposal scenarios. However, we also identify the ingrowth of a Tc colloidal phase that is stable under reducing conditions. The presence of ^{99}Tc in a mobile colloidal phase may lead to enhanced mobility of ^{99}Tc in the subsurface and clearly warrants further research. Overall, these data highlight the complex behavior of ^{99}Tc under repeated biogeochemical redox cycling with both Tc(IV) binding to Fe-phases, and Tc colloid formation impacting its potential fate and mobility. Furthermore, the processes explored in this study have clear implications when considering the eventual fate of other risk-driving, long-lived, redox-active radionuclides (e.g., Se, U, Np, and Pu) as they will undoubtedly undergo similar redox cycling processes.

MATERIALS AND METHODS

Sediment Collection. Sediment was collected from a well characterized⁴³ field site (Latitude $54^\circ 26' 30''$ N, Longitude $03^\circ 28' 09''$ W). The site is representative of the Quaternary unconsolidated alluvial floodplain deposits that underlie the Sellafield Ltd. site. After collection, the sediment was immediately hand-picked to remove larger debris and rock fragments, and then stored in a sealed sterile plastic container at 10°C in the dark until use, with experiments starting within 4 weeks of sample collection so as to minimize potential changes to the indigenous microbial community. The sediment characteristics (elemental composition, bulk mineralogy etc.) are summarized in the [Supporting Information \(SI\)](#).

Experimental Approach. We designed our experiments using Tc concentrations of 0.5 and $400\ \mu\text{M}$ to allow both straightforward radiometric counting of aqueous ^{99}Tc , and at $400\ \mu\text{M}$, to permit access to direct spectroscopic and electron microscopy analyses on samples thus allowing us to probe the molecular speciation and fate of the radionuclide. These Tc concentrations approximate to the wasteform environment and/or impacted environments in the immediate vicinity of nuclear facilities rather than typical sub nanomolar concentrations observed in far field impacted environments.⁷

Microcosm Preparation. Using aseptic technique, duplicate large volume microcosms were prepared with sediment ($100 \pm 1\ \text{g}$) and a synthetic groundwater representative of the Sellafield region ($1\ \text{L} \pm 1\ \text{mL}$; see the [SI](#) for the Groundwater composition) in 2 L bottles (Schott, Duran). The headspace was flushed with Ar gas and sealed with an airtight stopper. Prior to microcosm setup, the groundwater had been autoclaved to ensure sterility, flushed with Ar gas, and adjusted to $\text{pH} \sim 7.4$ using weak NaOH.⁴ Thereafter, $10\ \text{mM}$ acetate was added as an electron donor and $0.5\ \mu\text{M}$ ($30\ \text{Bq mL}^{-1}$) ammonium pertechnetate (CERCA, France) was spiked into the experiment. Microcosms were incubated in the dark at 21°C and sampled periodically. Control microcosms were prepared in triplicate and sterilized by autoclaving prior to the addition of TcO_4^- .

Geochemical Monitoring. Sediment slurry was sampled using aseptic technique under O_2 free conditions at each time point and the solids/supernatant were separated by centrifugation at $1.44 \times 10^4\ \text{g}$ for 10 min. Assuming a density of quartz ($2.65\ \text{g cm}^{-3}$) for any particles in solution, this centrifugation would sediment particles of $>450\ \text{nm}$ diameter, as predicted by Stokes' Law. For denser particle types, smaller particles would have been sedimented (e.g., $>290\ \text{nm}$ for magnetite ($5.15\ \text{g cm}^{-3}$)). To gauge the extent of microbially mediated Fe(III)-reduction/Fe(II) oxidation, the "bioavailable" $\text{Fe(II)}/\text{Fe(total)}$

Table 1. Details of EXAFS Fitting Parameters^a

sample	path Tc-	R (Å)	NoP	σ^2 (Å ²)	ΔE^0 (eV)	S_0^2	R-factor	Δ (%)
1st reduction	O	2.02 (1)	6*	0.005	0.80	0.95*	0.006	97.5
	Tc	2.55 (1)	1.36 (22)	0.003*				
	Fe	2.62 (4)	0.53 (30)	0.006*				
	O–Tc–O ^{ms}	4.05 (1)	6*	0.01				
1st oxidation	O	2.01 (1)	6*	0.005	0.31	0.95*	0.005	99.8
	Tc	2.56 (1)	1.36 (20)	0.003*				
	Fe	2.61 (2)	0.69 (28)	0.006*				
	O–Tc–O ^{ms}	4.03 (1)	6*	0.011				
2nd reduction	O	2.01 (1)	6*	0.005	0.73	0.95*	0.006	99.6
	Tc	2.56 (1)	1.17 (24)	0.003*				
	Fe	2.63 (3)	0.66 (30)	0.006*				
	O–Tc–O ^{ms}	4.04 (1)	6*	0.01				
2nd oxidation	O	2.01 (1)	6*	0.005	−0.3	0.95*	0.007	100.0
	Tc	2.56 (1)	0.84 (26)	0.003*				
	Fe	2.65 (4)	0.69 (27)	0.006*				
	O–Tc–O ^{ms}	4.04 (1)	6*	0.011				
3rd reduction	O	2.01 (1)	6*	0.005	0.3	0.95*	0.010	99.9
	Tc	2.55 (1)	0.72 (29)	0.003*				
	Fe	2.65 (5)	0.71 (31)	0.006*				
	O–Tc–O ^{ms}	4.04 (2)	6*	0.011				

^aR = atomic distance; NoP = number of paths; σ^2 (Å²) = Debye-Waller factor; ΔE^0 (eV) = energy shift from Fermi level; S_0^2 = amplitude factor; R-factor = normalized least squares residual. To determine change in Tc–Tc NoP and Tc–Fe NoP their σ^2 were fixed to 0.003 and 0.006, respectively. The amp (S_0^2) for the O shell was fixed. *Denotes fixed parameter, ^{ms} denotes linear multiple scattering path, and their R and σ^2 values were calculated as multiples of the single scattering path parameters. Δ confidence level (%) that the improvements in the fits with the “Tc–Fe” shell are significantly improved compared to the hydrous TcO₂ model alone (see SI) using the F-test.⁴⁶ Errors are presented in parentheses and are 1 σ of last decimal.

ratio in the sediment was determined by extraction of sediment pellet slurries with 0.5 M HCl for 1 h.⁴⁴ Fe(II) and Fe(total) was then measured in the supernatant (separated from the solids by centrifugation at 1.44×10^4 g for 10 min) using the ferrozine assay.⁴⁵ This method may not accurately extract all bioavailable Fe in sediment systems⁴⁶ but was used in this study to detail gross changes in “bioavailable” solid phase Fe speciation (i.e., provide an indicator as to the extent of Fe(III) reduction and Fe(II) oxidation during redox cycling experiments). At the start of the experiments the total amount of 0.5 N HCl extractable “bioavailable” Fe(III) was 5.30 ± 0.20 mmol kg^{−1}. Total ⁹⁹Tc concentrations in centrifuged (1.44×10^4 g for 10 min) solutions taken from sediment slurry samples were measured using liquid scintillation counting with appropriate standards and background subtraction such that the detection limit was 0.17 nM (~ 10 mBq ml^{−1}) (Wallac Quantulus 1220). The microcosms were also sampled at regular intervals for pH using calibrated electrodes and a Mettler Toledo digital meter. UV–vis spectroscopy methods were used to measure NO₂[−], Mn, and Fe in centrifuged (1.44×10^4 g for 10 min) solution samples^{45,47,48} using a Shimadzu UV-1800 UV spectrometer. Throughout, calibrations had typical regressions of $R^2 > 0.99$. Acetate, SO₄^{2−} and NO₃[−] were measured from the same centrifuged solutions using a Dionex DX120 ion chromatograph. Eh was not routinely measured during experimentation due to sample volume constraints; however, it was measured from centrifuged solution samples taken at the end of the first and third reduction and oxidation periods. The samples used for Eh measurement had been stored at -80 °C under Ar prior

to analysis. They were then defrosted and centrifuged (1.44×10^4 g for 10 min) and the solution Eh was measured using a calibrated electrode. Microcosms were also weighed between sampling points to monitor for evaporation, which was found to be negligible.

Redox Cycling Experiments. After initial reduction (signified by the 0.5 M HCl-extractable Fe(II)/Fe(total) ratio reaching $\sim 100\%$ Fe(II)), the microcosms were oxidized with air (containing O₂ as the oxidant). Here, the headspace was flushed regularly (every 1–3 days) with 1 L of gas passed through a 0.22 μ m filter (80:20 air: CO₂ mix), to maintain a near neutral pH. This gas treatment continued until the proportion of 0.5 M HCl-extractable Fe(II) in the sediment had significantly reduced and stabilized between sampling points. At this point, acetate (5 mM) was added as an electron donor to stimulate anaerobic metabolism and the headspace was flushed with 0.22 μ m filtered Ar. The microcosms were then incubated until the proportion of 0.5 M HCl-extractable Fe(II) had stabilized between sampling points at between 95–100% of the total extractable Fe. In total, three full reduction/oxidation events were sampled. Decisions on when to start/stop oxidation treatments were not based on achieving steady state for Tc concentrations in solution.

Acid Extraction of Sediment Samples. Acid extractions were performed on select sediment samples to assess any relationship between Fe and Tc in redox-cycled sediments. Here, sediments were sequentially leached with increased molarity HCl for 30 min (unless stated), in the order: 0.001, 0.01, 0.1, 1, 2, 3, 4, and 6 M, 6 M for 24 h, 6 M for 4 days and,

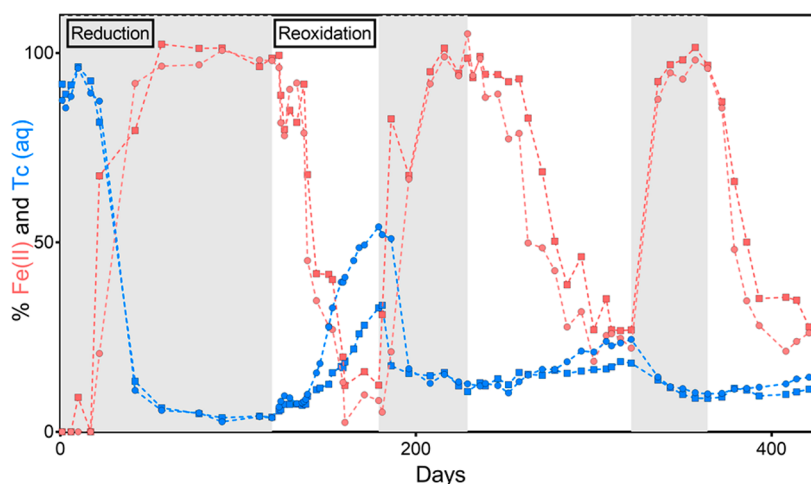


Figure 1. Tc remaining in centrifuged (1.44×10^4 g, 10 min) solutions (blue dashed line) (expressed as a percentage of the original ^{99}Tc spike ($0.5 \mu\text{M}$)) and the % of 0.5 N HCl extractable Fe(total) present as Fe(II) in sediments (extracted by 0.5 N HCl for 1 h; total bioavailable Fe(III) = $5.30 \text{ mmol kg}^{-1}$) (red dotted line), for duplicate microcosms (microcosm 1: circles, microcosm 2: squares). Shaded areas denote “reduction events”, where microcosms were amended with acetate and kept anaerobic. Unshaded areas denote “oxidation events”, where microcosms were oxidized with air (O_2). The analytical error associated with ^{99}Tc measurements was low ($<1\%$ of the solution ^{99}Tc concentration); the limit of detection for Tc was 0.17 nM . The reproducibility of the 0.5 N Fe extraction was $\sim \pm 10\%$ based on repeated measurement ($n = 3$) of the same sample.

finally, in heated *aqua regia* for 4 h (modified from a previously published method;⁴⁹). The centrifuged (1.44×10^4 g, 10 min) extracts were analyzed for ^{99}Tc and Fe using liquid scintillation counting and the ferrozine method, respectively.

Ultrafiltration of Groundwater Samples. Ultrafiltration of solutions was performed at the end of the third reduction and oxidation periods using Nanosep Centrifugal Filters (Pall Corp.; centrifuged at 1.44×10^4 g, 10 min) to explore the colloidal/solution distribution of Tc. Two filter sizes were used; $0.2 \mu\text{m}$ and 10 kDa , the latter having a geometric mean particle size of $\sim 3 \text{ nm}$ for a globular particle.⁵⁰ The filters were pre-equilibrated with sample solution prior to use. Technetium concentrations in the filtrate were measured using liquid scintillation counting.

X-ray Absorption Spectroscopy. Higher activity parallel samples (2.5 g sediment in 25 mL groundwater, at a concentration of $400 \mu\text{M}$ Tc(VII)) were incubated at the elevated concentrations required for XAS in order to measure ^{99}Tc speciation in sediment and solution during redox cycling. Sediment was sampled (as previously described) at the end of reduction periods (days = 266, 325, 380) and at the end of two oxidation periods (days = 310 and 372). For oxidation time-points, sediment samples were rinsed with thoroughly deoxygenated water, under O_2 free conditions prior to analysis.³² Centrifuged (1.44×10^4 g, 10 min) solution samples for XANES measurement were taken at the end of the second oxidation period and third reduction period. All XAS samples were prepared, packed into airtight, triple contained sample cells, and stored under O_2 free conditions. XANES and EXAFS spectra were collected on beamline B18 of Diamond Light Source from the ^{99}Tc K-edge in fluorescence mode at 78 K . Spectra were collected with either a 9 or 36 element Ge detector. XANES data underwent linear combination fitting (LCF) between Tc(IV) and Tc(VII) (as ammonium pertechnetate) standards⁵¹ to provide an indication of the percentage distribution of Tc(IV) and Tc(VII) in the samples. EXAFS data were fitted using Athena and Artemis software packages and using relevant information from the literature.⁵² The statistical viability of additional shells was assessed using

the F-test.^{29,53} To determine the trend in the Tc–Tc coordination number, the fitting procedure was constrained to fixed values for the passive electron reduction factor (S_0^2) and Debye–Waller factor (σ^2) (Table 1).

Transmission Electron Microscopy. Transmission electron microscope (TEM) grids were prepared under O_2 free conditions with the groundwater from high activity (XAS) microcosms taken at the end of the third reduction period. Prior to preparation, the groundwater was centrifuged at 1.44×10^4 g for 10 min and a small sample of the supernatant was then pipetted on the TEM grid under O_2 free conditions. Images were taken using a FEI Tecnai TF20 field emission gun TEM and the particle composition on the TEM grids was analyzed with an Oxford Instruments INCA 350 EDX. Mineral composition was quantified using Selected Area Electron Diffraction (SAED).

16S rRNA Gene Amplicon Pyrosequencing and Data Analysis. DNA from both microcosms was isolated from $\sim 0.2 \text{ g}$ of sample from the starting sediment (day 0), sediments after reduction events (days = 119, 229, 372), and sediment after oxidation events (days = 179, 321, 428) using the MoBio PowerSoil DNA Isolation Kit (MoBio Laboratories Inc.). PCR for amplicon pyrosequencing was performed using tagged fusion bacterial primers 27F⁵⁴ and 338R,⁵⁵ as described previously.⁵⁶

Pyrosequencing was performed using a Roche 454 Life Sciences GS Junior system. The 454 pyrosequencing reads were analyzed using Qiime 1.8.0 release,⁵⁷ and denoising and chimera removal was performed in Qiime during operational taxonomic units (OTU) picking (at 97% sequence similarity) with usearch.⁵⁸ Taxonomic classification of all reads was performed in Qiime using the *ucrust* method against the Greengenes database, release 8/2013.⁵⁸

Nonmetric multidimensional scaling (NMDS) was performed on Primer v6 (PRIMER-E Ltd., UK), based on a Bray–Curtis similarity matrix⁵⁹ that was calculated using square root transformed relative (%) abundances of all the identified OTUs of this study.

■ RESULTS AND DISCUSSION

Redox Cycling and Changes in Solution Technetium Concentrations. Sediment microcosms representative of the Sellafield subsurface were run in duplicate to explore the fate of Tc during three reduction/oxidation cycles over ~1.5 years. The sediment microcosms (1:10, solid: solution ratio; volume 1 L) were amended with an electron donor (10 mM acetate), spiked with Tc(VII) (0.5 μ M) and incubated anaerobically until Fe(III)- and SO_4^{2-} -reducing conditions had developed. Periodically, microcosms were then oxidized with air (O_2) and then reduced through further addition of acetate, which promoted microbially mediated reduction following the normal sequence of terminal electron accepting processes (Figure 1 and SI Figure 1). At the end of periods of reduction, the Eh of the microcosms was ≤ -140 mV; at the end of periods of oxidation the Eh was ≥ 187 mV. These values are similar to past Eh measurements in Sellafield sediment systems.^{8,39} In sterile (oxic) control experiments, redox conditions remained unchanged and Tc remained in the centrifuged solution (1.44×10^4 g, 10 min; referred to as “solution” throughout the manuscript unless otherwise stated) as expected for Tc(VII) under oxic/denitrifying conditions (SI Figure 2^{9–11}). In the microbially active experiments, Tc remained in solution until ~20 days; thereafter, ⁹⁹Tc was largely scavenged to solids concomitant with Fe(III)-reduction (~94% of added Tc removed from solution between days 22 and 57, Figure 1). Here, scavenging of Tc was likely dominated by abiotic reduction of Tc(VII) through reaction with biogenic Fe(II).^{10,24} After 119 days, the proportion of 0.5 N HCl-extractable Fe(total) present as Fe(II) in solids (herein referred to as extractable Fe(II)) was $98.3 \pm 0.5\%$ and the Tc concentration in solution had stabilized ($96.2 \pm 1.1\%$ of added Tc removed from solution, 18.2 ± 0.6 nM remaining in solution) near the theoretical solubility of hydrous TcO_2 (~10 nM; 35). At this point, oxidation of the microcosms was initiated by flushing the headspace with 0.22 μ m filtered air (O_2 being the oxidant) every 1–3 days for 60 days. This resulted in the reoxidation of extractable Fe(II), and after 60 days of air amendment, the extractable Fe(II) in solids had stabilized at $10.3 \pm 2.9\%$ (Figure 1). The oxidation of Fe(II) was accompanied by significant, yet incomplete Tc remobilization to solution with $43 \pm 15\%$ of the initial Tc spike remobilized from the sediment after 60 days of oxidation. This level of Fe(II)-reoxidation and Tc remobilization during air oxidation is in agreement with past work completed over similar time scales.^{32,33,37,39,41}

After oxidation had stabilized with a measured sediment extractable Fe(II) proportion of $10.3 \pm 2.9\%$ (day 179, Figure 1), bioreduction was again stimulated with the addition of 5 mM acetate and with purging of the headspace with 0.22 μ m filtered Ar. Here, microbially mediated Fe(III)-reducing conditions rapidly re-established alongside removal of Tc from solution. During this second period of bioreduction, extractable Fe(II) in solids reached ~100% within 35 days (Figure 1), and removal of SO_4^{2-} from solution also occurred, suggesting the onset of microbially mediated SO_4^{2-} reduction (see SI Figure 1). Interestingly, despite these strongly reducing conditions, a greater proportion of Tc remained in solution when compared to the first reduction period ($11.6 \pm 1.4\%$ compared to $3.8 \pm 0.1\%$; Figure 1). After 229 days, the experiments were again reoxidized through repeated (every 1–3 days) purging of the headspace with air for a period of 92

days. During the second period of reoxidation, a larger proportion of extractable Fe(II) remained stable in the solids ($24.5 \pm 3.4\%$ by day 321), suggesting that there was formation of an Fe(II)-bearing phase that was more resistant to oxidation such as magnetite. In addition, less Tc was remobilized into solution during the second reoxidation ($43 \pm 15\%$ in the first oxidation period vs $21 \pm 4\%$ in the second; Figure 1). Finally, a third period of bioreduction (addition of 5 mM acetate, headspace purging with filtered Ar, and anaerobic incubation for 43 days) and oxidation (controlled exposure to filtered air for 64 days) was completed. At the end of the third period of reduction, $96.4 \pm 0.6\%$ of the extractable Fe was present as Fe(II) and $9.4 \pm 0.9\%$ of the added Tc remained in solution. After subsequent reoxidation, the extractable sediment Fe(II) content again stabilized at around $26.4 \pm 1.1\%$ and Tc in solution increased to $13.1 \pm 2.6\%$. Overall, during redox cycling, a significant fraction (~25%) of extractable Fe(II) became resistant to oxidation, potentially due to ingrowth of recalcitrant Fe(II) bearing phase(s). By contrast, the total pool of 0.5 N HCl extractable Fe remained relatively constant during redox cycling (see the SI and SI Table 1). For Tc, there was a significant increase in Tc measured in solution between the first ($3.8 \pm 0.1\%$) and third reduction periods ($9.4 \pm 0.8\%$). In contrast, during oxidation, the sediment associated Tc became progressively more resistant to remobilization ($43 \pm 15\%$ remobilized to solution during the first oxidation compared to $13.1 \pm 2.6\%$ in the third). This suggests that with repeated redox cycling, a significant fraction of sediment associated Tc becomes more strongly resistant to remobilization while at the same time, surprisingly, the concentration of Tc that remains in solution under reducing conditions increases.

Technetium Extractability and Speciation on Sediments during Redox Cycling. In order to further investigate the relationship between Tc and Fe in the sediment we used acid extractions on solid phases at the end-points of the first and third reoxidation periods (days 179 and 428, Figure 1). The samples showed the same general Tc and Fe dissolution trends in the acid extractions (Figure 2; concentration data presented in SI Table 2), but the Tc and Fe from the first reoxidation period samples were always more susceptible to acid leaching than those in samples from the third reoxidation period. This suggests that there is an increased recalcitrance to remobilization for Tc with repeated redox cycling. Further, in both treatments, after sequential leaching for 30 min with HCl concentrations ranging from 0.001 to 6 M, the slope of Tc vs Fe dissolution was always >1 . This indicated that most of the Tc was heterogeneously distributed on the mineral surfaces, as also indicated by an enhanced release of available Tc (~40–80%) compared to Fe (<20%) in the first and third reoxidation samples, respectively. However, the Tc became more resistant to acid leaching with repeated redox cycling (Figure 2), with $25 \pm 5\%$ (first reoxidation) and $52 \pm 8\%$ (third reoxidation) resistant to leaching with 6 M HCl for 30 min. Indeed, an aggressive, oxidizing *aqua regia* leach was required for full dissolution of the Tc from the solids (Figure 2; SI Table 2).

The marked decrease in Tc susceptibility toward remobilization (Figure 1) and the increased recalcitrance of Tc to acid leaching (Figure 2) between the first and third periods of reoxidation confirm that the Tc became more strongly bonded within the sediment as a result of repeated redox cycling. This increased recalcitrance to oxidative remobilization may be associated with enhanced Tc association with Fe in the sediments. Indeed, Fe(II) has been suggested to act as a buffer

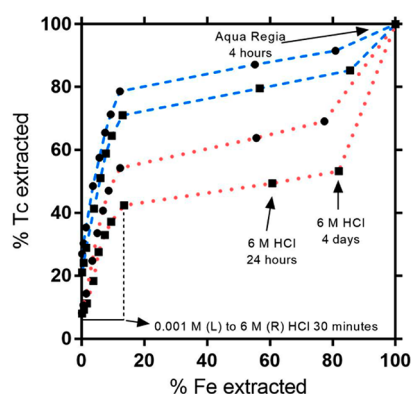


Figure 2. Cumulative dissolution of Tc versus Fe from select sediment samples. Samples are the 1st oxidation (day 179, blue dashed lines, microcosm 1: circles, microcosm 2: squares) and 3rd oxidation (day 428, red dotted lines, microcosm 1: circles, microcosm 2: squares). The $[H^+]$ in the 30 min extraction was increased from 0.001 to 0.01 M, 0.1 M, 1 M, 2 M, 3 M, 4 M, and, 6 M HCl, with the final three extractions being 6 M HCl for 24 h, 6 M HCl for 4 days, and heated *aqua regia* for 4 h. The axes are specific to the concentration of Tc or Fe extracted per sample, with 100% being the maximum amount of Tc or Fe extracted from that sample (concentration data are presented in SI Table 2). The final *aqua regia* digest extracted close to all of the available sediment associated Tc per sample, and ~25% of the total sediment Fe as measured by XRF (3.71 wt %).

for Tc(IV) oxidative remobilization in sediment and clay systems^{28,32} and Tc(IV) may have become physically protected from remobilization through bonding to and/or armoring by Fe(II)/Fe(III) oxidation products, or via incorporation into Fe minerals.^{26,29,31,32} We also note that past work⁶⁰ has identified possible oxidation of Tc(IV) by Fe(III) at acidic pH, but in this study Fe(III) mediated Tc(IV) oxidation appears to be of minimal importance. This is demonstrated by the fact that the samples that underwent three redox cycles had a higher proportion of Fe(III) (Figure 1; SI Table 2) yet showed stronger Tc retention (Figure 2) when compared to the samples that underwent one redox cycle (Figure 1; SI Table 2).

To further assess potential changes in the Tc speciation and coordination environment in these sediment systems, we undertook X-ray Absorption Spectroscopy (XAS) on a select

group of parallel experiments run at the higher (~0.4 mM) Tc concentrations necessary for detection by XAS. Data were collected at the end of the first, second, and third reduction periods, and at the end of the first and second oxidation periods (see SI). All the Tc K-edge X-ray Absorption Near Edge (XANES) spectra aligned with a Tc(IV) reference sample, confirming that there was no significant Tc(VII) in any of the samples, regardless of whether they were taken from reduced or reoxidized sediments (Figure 3A). Further, none of the Tc XANES spectra had evidence of a double peak near the white line which can be indicative of Tc incorporation in Fe (oxyhydr)oxides.^{29,61}

Analysis of the Extended X-ray Absorption Fine Structure (EXAFS) Fourier transform (FT) data for all samples showed two shells of backscatters were present; the first shell at ~2.0 Å and the second shell at ~2.5 Å (Figure 3B and C). There was also a consistent trend in decreasing FT amplitude for the ~2.5 Å backscattering shell across the redox cycles (Figure 3B), suggesting a reduced occupancy for backscatters in this shell with repeated redox cycling. Broadly, these data are similar to the model for Tc(IV)O₂·xH₂O chains¹³ and as seen in several studies where Tc is present in environmental systems (e.g.,^{20,26,32,62}). The Tc(IV)O₂·xH₂O coordination environment was therefore used as the starting point for modeling. The data were initially fitted with fixed S_0^2 and σ^2 values and with 6 O backscatters at 2.0 Å, with the best fits realized by optimizing the number of Tc backscatters at ~2.5 Å (see SI Table 3). This modeling approach suggested that the average coordination number of Tc in the backscattering shell at ~2.5 Å decreased during the redox cycles, starting at 1.2 ± 0.1 Tc for the first reduction and falling to 0.7 ± 0.1 Tc by the third reduction (see SI Table 3). Interestingly, the observed decrease in the Tc coordination number at ~2.5 Å implies that a shortening of the octahedral chain length in Tc(IV)O₂·xH₂O chains¹³ is occurring during the redox cycles. In the “hydrous TcO₂” model, a Tc coordination number of two represents an infinite chain as each central Tc atom is neighbored by two Tc atoms, a coordination number of one represents a TcO₂ dimer, and a coordination number of zero represents a TcO₂ monomer. These fitting results (SI Figure 3 and Table 3) suggest that dimers dominate during the first cycle of reduction and oxidation, tending toward monomers with repeated redox

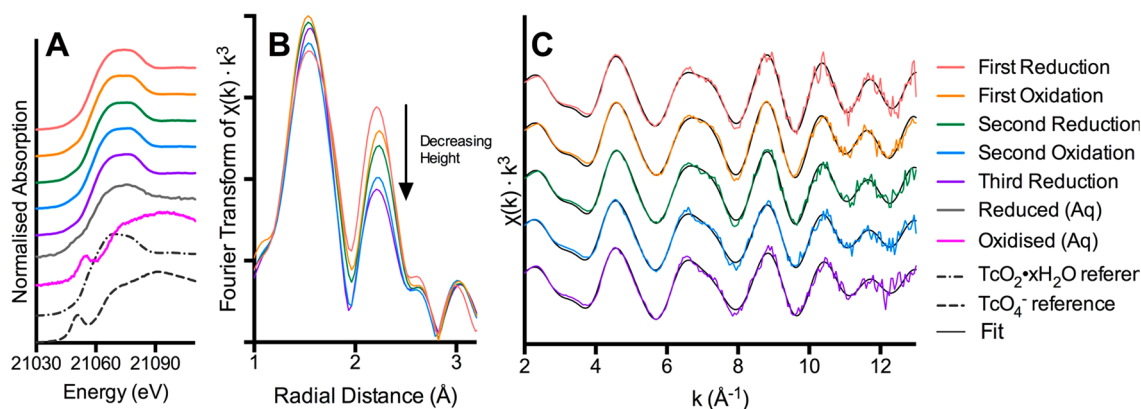


Figure 3. XAS data for ⁹⁹Tc in reduced and oxidized sediment samples, and solutions separated from the sediment by centrifugation. Panels A; Tc K-edge XANES spectra for reduced and oxidized sediment, solution samples from reduction and oxidation samples, and reference spectra (from Hess et al., 2004; 51). B; Nonphase shifted Fourier transform of k^3 -weighted EXAFS spectra calculated over a k range of 3.5–13 Å⁻¹, with their best fit models (including “short Fe” interaction), and C; k^3 -weighted EXAFS spectra (colored lines) with their best-fit models. For comparison, representative EXAFS spectra for TcO₂ and Tc(IV) associated with magnetite can be found in refs 59 and 29.

cycling. However, the EXAFS between 10.5–13 K in the samples also had features that were not fully resolved with this model alone. Reflecting this, Tc binding within sediments via an edge-sharing “short Fe” octahedron at ~ 2.6 Å was also considered in fitting (Figure 3B and C; Table 1). This has been attributed to attachment of the TcO_2 chains at Fe–O surfaces.^{26,31,32,63} In all cases, this additional contribution from a “short Fe” backscatterer at ~ 2.6 Å improved the best fits (Table 1). As expected, these constrained fits also showed a trend of decreasing Tc coordination number at ~ 2.5 Å with progressive redox cycles, from 1.4 ± 0.2 in the first reduction to 0.7 ± 0.3 in the third reduction, and consistent with shortening of the TcO_2 chain. At the same time, there was an increase in both the proportion of Tc–Fe bonding at ~ 2.6 Å, and the confidence that the additional “short Fe” improved the fits with repeated redox cycling (Table 1, SI Table 3).^{20,29,64} This confirms the samples can be consistently fitted with a physical model showing a relative decrease in the TcO_2 chain length, coupled with a proportional increase in Tc–Fe bonding via attachment to surface Fe–O sites throughout progressive redox cycles. Importantly, the molecular scale changes in Tc(IV) binding observed in these systems offer key insights into bulk Tc behavior in the sediments. Specifically, with repeated redox cycling, we see evidence for hydrous TcO_2 chain shortening and for the TcO_2 chains to become increasingly associated with edge-sharing Fe–O. These molecular scale observations then manifest as a decrease in oxidative remobilization of Tc with repeated redox cycles, and are supported by the observation that bonding of short chain hydrous TcO_2 protects or renders Tc(IV) less susceptible to oxidation.³² In summary, our data suggest that repeated redox cycling produces a shorter TcO_2 chain that is strongly bonded to Fe–O in sediments. In turn, and in agreement with the sequential extraction data, these molecular scale observations are linked to the increased recalcitrance of sediment-associated Tc(IV) to remobilization as seen with repeated redox cycling.

Technetium Colloid Formation. While the sediment associated Tc appeared to become more recalcitrant to oxidative remobilization, the concentration of Tc remaining in the aqueous phase increased during successive redox cycles (Figure 1; $\sim 4\%$ after the first reduction cycle to $\sim 13\%$ at the end of the third oxidation cycle). Past work has highlighted the presence of a “soluble” Tc phase under reducing conditions with both colloidal Tc(IV)^{16,65,66} or Tc(IV) carbonate complexes⁶⁷ reported or inferred. To examine this further, centrifugation and filtration of the separated aqueous phase was used to assess the presence of Tc colloids in these systems. In the lower concentration microcosms ($0.5 \mu\text{M}$ starting concentration of Tc) after the third oxidation event, practically all of the Tc in the aqueous phase ($98.2 \pm 5.5\%$) passed through a 10 kDa (~ 3 nm) filter. In contrast, after the third reduction event, $94.5 \pm 3.0\%$ of the Tc passed through a 200 nm filter but only $65.9 \pm 0.1\%$ through a 10 kDa (~ 3 nm) filter, suggesting that approximately one-third of the Tc was present as a colloidal fraction of between ~ 3 –200 nm at the end of the third reduction cycle. This is consistent with past work that has documented elevated Tc(IV) concentrations in the aqueous phase above the theoretical solubility threshold for hydrous TcO_2 .^{8,16,66,68} The remaining $\sim 2/3$ of the Tc was presumably present as Tc(IV) aquo-ions or complexed Tc(IV), or colloidal nanoparticles with a size of <3 nm. In the higher Tc concentration samples run for XAS, we were able to analyze the solution from the third reducing cycle using both XAS and

TEM approaches. The TEM analyses showed clear evidence for colloids with Fe-rich aggregates of 20–100 nm diameter clearly visible. These colloids had clear Tc and Fe signals in EDX spectra (Figure 4). Indeed, the aggregates seemed to be

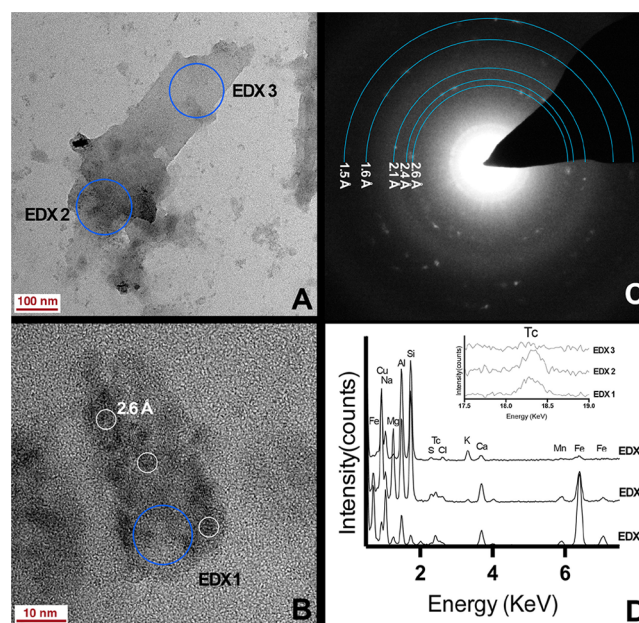


Figure 4. Panels A and B are TEM images of colloids at the end of the third reduction period; selected areas with lattice spacings of 2.6 Å are signified by white circles; blue circles correspond to the areas where energy dispersive X-ray (EDX) spectra were collected. Panel C shows selected area electron diffraction of the particle in panel B, light blue lines show measured d -spacings (Å) of the diffraction spectrum. Panel D shows EDX data collected from Panels A and B.

composed of 5–10 nm nanoparticles and SAED (Figure 4) on the TEM samples, were consistent with magnetite formation, a mineral formed both during microbially mediated Fe(III) reduction⁶⁹ and when Tc(VII) reacts with aqueous Fe(II).³¹ We were also able to obtain Tc K-edge XANES data for the third reduction and second reoxidation solution cycle samples despite their low concentration (Figure 3A). Here, the XANES spectrum for the reduced sample was clearly Tc(IV)-like in structure and linear combination fitting between Tc(IV) and Tc(VII) standards confirmed that Tc(IV) dominated the sample (see SI Table 4). For comparison, XANES data from the solution sample taken during oxidizing conditions clearly resembled Tc(VII) (Figure 3A;⁵¹).

Confirmation by ultrafiltration and TEM of iron (oxyhydr)-oxide colloid associated Tc, which appears to be present as Tc(IV) under reducing conditions in these experiments, has significant implications for the mobility of Tc as colloids can greatly increase radionuclide migration in the subsurface.^{70,71} Indeed, the significantly enhanced concentration of colloid bound Tc(IV) in these experiments under reducing conditions, when compared to the expected solubility for TcO_2 (as high as $\sim 33 \mu\text{M}$ in our higher activity experiments compared to a predicted solubility of ~ 10 nM),³⁵ highlights the need for further work on these phases, for example, understanding Tc(IV)-Fe (oxyhydr)oxide colloid formation pathways, and stability in environmental systems.^{29,31} Also, this study did not determine changes in dissolved organics (beyond acetate, SI Figure 1) during redox cycling. As dissolved organics have been

implicated in Tc behavior in environmental systems (including Tc(IV)-colloids) (e.g.,^{72,73}), their role in longer-term redox cycling studies warrants further investigation.

Microbial Community Response to Redox Cycling.

16S rRNA gene amplicon pyrosequencing was also used to assess the microbial community response to repeated redox cycles, in both microcosms. Analysis was possible in the low Tc concentration experiments for the starting sediment and at the end points of each reduction and oxidation event (Figure 1; days = 0, 119, 179, 229, 321, 372, 428). Overall, the results show a diverse microbial community in all sequenced samples (see SI Figure 4 and Table 6), and sequences related to known Fe(III)- and Tc(VII)-reducing genera, most notably the deltaproteobacteria *Geobacter*²⁴ and *Anaeromyxobacter*⁷⁴ were detected. Additionally, there were broader responses in the microbial community compositions to oxidizing or reducing conditions. For example, organisms affiliated with obligate anaerobic Firmicutes of the class Clostridia increased in relative abundance after reduction events, and were less abundant in the community profiles after every oxidation event (SI Figure 4 and Table 6). The opposite correlation was noted within the betaproteobacteria, where close affiliates of the family Comamonadaceae, the genus *Azospira*, and the order Burkholderiales (which all exhibit aerobic/microaerophilic metabolisms) generally increased in abundance during oxidizing cycles. Similarities between samples based on microbial community composition were explored further using nonmetric multidimensional scaling (NMDS)⁵⁹ and this showed that oxidation periods had more of an impact on sediment microbial community structure when compared to reduction periods (SI Figure 5).

Implications for Technetium Behavior in the Environment. Here, for the first time we have shown that during repeated redox cycles, Tc(IV) becomes increasingly strongly bound to sediments. Spectroscopic analysis confirmed a correlation between Tc(IV) recalcitrance to reoxidation and molecular scale changes in Tc(IV) speciation. Repeated redox cycling prompted a shortening of TcO₂ chains and an increase in associations with Fe–O edge sites in the sediments. At the same time, chemical extractions confirmed an increase in Tc strongly associated with the sediments with successive redox cycles and that Fe(II) may also play a role in buffering sediment associated Tc(IV) oxidative remobilization. This suggests that the current literature which only addresses single oxidation events, may overestimate the influence of oxidation on long-term Tc mobility. By contrast, analysis of solutions showed an ~3000× enhancement in Tc concentrations after several redox cycles compared to those predicted from thermodynamic calculations, with analysis (ultrafiltration and TEM) documenting the presence of an Fe bearing colloidal phase in solution that also contains Tc under anoxic conditions. This has significant and, as yet, largely unexplored implications for Tc mobility in reducing environments.

Overall, this study has revealed a number of unique observations on the environmental behavior of technetium in long duration redox cycling experiments and emphasizes the value of this longer-term approach in refining models of the environmental behavior of this radionuclide. Indeed, radionuclide fate upon multiple redox cycling seems to be a largely unexplored yet critical variable in predicting the longer-term environmental behavior of Tc, and by implication may be significant for other long-lived radioactive contaminants such as Se, Np, U, and Pu. Going forward, it is important to understand

whether these radionuclides respond to ephemeral redox conditions; further, experimentation needs to move from static microcosms systems to more dynamic systems (e.g., columns/lysimeters) (e.g., ref 75).

■ ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acs.est.7b02426.

Supplementary sections presenting ancillary data, extra information on methods, and Tables and Figures for control experiments, additional geochemical indicators, sediment characterization data, sequential extraction concentration data, alternative EXAFS fits, XANES linear combination fitting, and microbial community data (PDF)

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Notes

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